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Daniel Jacob, atmospheric chemistry,

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Introduction to Atmospheric Chemistry is a concise, clear review of the fundamental aspects of atmospheric chemistry. In ten succinct chapters, it reviews our basic understanding of the chemistry of the Earth's atmosphere and discusses current environmental issues, including air pollution, acid rain, the ozone hole, and global change.

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Atmospheric chemistry is one of the fastest growing fields in the earth sciences. Until now, however, there has been no book designed to help students capture the essence of the subject in a brief course of study. Daniel Jacob, a leading researcher and teacher in the field, addresses that problem by presenting the first textbook on atmospheric chemistry for a one-semester course.

~~Introduction to Atmospheric Chemistry on
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Introduction to Atmospheric Chemistry (Princeton University Press, 1999). They are arranged following the different chapters of the book. In recent years I have added to my course lectures a chapter 14, 'Aerosol Chemistry' and a chapter 15, 'Mercury in the Environment'. I have included problems to support these chapters.

~~INTRODUCTION TO ATMOSPHERIC CHEMISTRY~~

Atmospheric chemistry spans from the fundamental understanding of reactions to accurately measuring reaction rate and photochemical data for gas-phase processes. Papers by Smith, Troe, Golden, and Donahue address the fundamental kinetics issues.

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Atmospheric chemistry is a branch of atmospheric science in which the chemistry of the Earth's atmosphere and that of other planets is studied. It is a multidisciplinary approach of research and draws on environmental chemistry, physics, meteorology, computer modeling, oceanography, geology and volcanology and other disciplines. Research is increasingly connected with other areas of study such as climatology. The composition and chemistry of the Earth's atmosphere is of importance for several re

~~Atmospheric chemistry — Wikipedia~~

1. The saturation vapor pressure of water at 293 K is $P_{H_2O, SAT} = 23$ hPa. At sunset the air is at 50% relative humidity, therefore $P_{H_2O} = 11.5$ hPa. The dew point corresponding to this water vapor pressure is 282 K. The air must cool to 282 K in order for fog to form. 2.

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Atmospheric Chemistry CHEM-5151 / ATOC-5151
Spring 2005 Prof. Jose-Luis Jimenez Lecture 1: Introduction to Atmospheric Chemistry
Outline of Lecture 1 • Importance of atmospheric chemistry • Atmospheric composition: big picture, units • Atmospheric structure – Pressure profile – Temperature profile – Spatial and temporal scales ...

~~Lecture 1: Introduction to Atmospheric Chemistry~~

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~~Introduction to atmospheric chemistry—
University of ...~~

Introduction to Atmospheric Chemistry. Daniel J. Jacob. Princeton University Press, Dec 21, 1999 - Science - 280 pages. 0 Reviews.

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The questions and problems presented here are intended to supplement my book Introduction to Atmospheric Chemistry (Princeton University Press, 1999). They are arranged following the different chapters of the book.

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